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SWBAT

determine the mass of
one substance involved
in a chemical reaction
if given the mass of
any one of the other
substances involved in

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the reaction. in other
words . . .

Assessment Answers

Slide 1

Solutions Manual
Chemistry: Matter and
Change • Chapter 8
121 Section 8.1 The
Covalent Bond pages
240–247 Practice
Problems page 244
Draw the Lewis
structure for each
molecule. 1. PH₃ H₂ H₂
H— H H P respectively,
for single, double, and
triple P — — 2. H₂ S H

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H H — H S S — 3. HCl H
Cl H — Cl 4. CCl 4 Cl Cl
Cl C C — —

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211 CHAPTER 11
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Section 11.1

Assessment page 372

5. Compare the mass of the reactants and the mass of the products in a chemical reaction, and explain how these masses are related. The coefficients in the balanced equation indicate the molar relationship between each

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Stoichiometry

Mass Finals for CHEM
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1-50 COUNT ONE
POINT EACH. 100 d) 10
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metric prefix kilo
means: 1,000 1000
—442.70

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Chemistry 121
Experiment 1 Report
Template Section Date:
Lab Partner:
DISCUSSION

QUESTIONS I. Explain

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the following terms:
accuracy of data,
precision of data, and
instrumental precision.

2. The density of a substance is considered an inherent property, unique to that substance. Is this property constant under all conditions? Explain.

3.

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Chemistry Section
Review 6.3 - 6.4
Multiple Choice Identify
the choice that best

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completes the statement or answers the question. Put the LETTER of the correct answer in the blank.

___ 1. The chemical formula for an ionic compound represents the a. number of atoms in each molecule. b. number of ions in each molecule.

Chemistry Section Review 6.3 - 6

Chapter 12 (pages 320
- 321) - Answer

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Section

questions - 1-9, 14.

End-of-Chemistry

Assessment Review. To

help you prepare for

the End-of-Chemistry

Assessment, review

the following at the

end of each chapter:

Chapter Highlights.

Vocabulary . Answer

Chapter Review

questions (answers will

be posted on my

website next week)

End-of-Chemistry Assessment Review

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